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P3-15 Assuming the reactions to be elementary: 2 Anthracene \rightarrow Dimer 2 () Dimer A Anthracene $C C r k C K$ where, $K c = k + /k-$
Similarly for the second reaction: 3-13 Norbornadiene Quadricyclane () Quadricyclane Norbornadiene Norbornadiene $C C r k C K$ where, $K c = k + /k-$
P3-16 The mistakes are as follows: 1. For any reaction, the rate law cannot be written on the basis of the stoichiometric equation.

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Stoichiometry. We shall set up Stoichiometric Tables using A as our basis of calculation in the following reaction. We will use the stoichiometric tables to express the concentration as a function of conversion. We will combine $C_i = f(X)$ with the appropriate rate law to obtain $-r_A = f(X)$.

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