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## **STUDY GUIDE FOR CONTENT MASTERY Section 12.2**

***Stoichiometric Calculations In your textbook, read about mole-to-mole conversion. Read the following passage and then solve the problems. In the equation that follows each problem, write in the space provided the mole ratio that can be used to solve the problem.***

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***Section 11.1 What is stoichiometry? In your textbook, read about stoichiometry and the balanced equation. For each statement below, write true or false. \_\_\_ true \_\_\_ 1. The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry.***

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### **[Section 12 1 What Is Stoichiometry Study Guide For Content ...](#)**

**STUDY GUIDE FOR CONTENT MASTERY The Periodic Table and Periodic Law Section 6.1 Development of the Modern Periodic Table In your textbook, reads about the history of the periodic table's development. Use each of the terms below just once to complete the passage. nine eight accepted octaves elements protons atomic mass properties periodic law**

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**Stoichiometry Study Guide KEY Chemistry RHS Mr. Moss STUDY GUIDE FOR CONTENT MASTERY Section 12.2 Stoichiometric Calculations In your textbook, read about mole-to-mole conversion. ... Chapter 12 Stoichiometry . In the reaction represented by the equation  $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$ , how many grams of hydrogen are produced if 120. g of Na and 80.0 g  $\text{H}_2\text{O}$  are used?**  
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**Stoichiometry Exam Study Guide. Chemistry GT. Don't forget to bring a calculator to the exam! Periodic table and Pink Sheets will be provided. Topics: Be able to complete mass - particles, mole-mole, mass-mass, mass-vol, vol-vol, and particles-particles problems. You should also be able to complete variations of the same types of problem.**

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**This first, short introductory SparkNote to Stoichiometry is meant to give you a very basic and general understanding of what stoichiometry is. It also introduces the most important concept of stoichiometry: the mole ...**

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**Stoichiometry is the tool for answering these questions. Stoichiometry The study of quantitative relationships between the amounts of reactants used and amounts of products formed by a chemical reaction is called stoichiometry. Stoichiometry is based on the law of conservation of mass. Recall from Chapter 3 that the law states that**

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**From a general summary to chapter summaries to explanations of famous quotes, the SparkNotes Stoichiometric Calculations Study Guide has everything you need to ace quizzes, tests, and essays.**

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